## *Lel<sub>2</sub>ob* and *Ob<sub>2</sub>lel* Cage Complexes Based on $[Co(pn)_3]^{3+}$ (pn = Propane-1,2-diamine): Synthesis, Resolution, and Tentative Identity of Their *Fac* and *Mer* Geometrical Isomers

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Encapsulation of  $[Co(R/S-pn)_3]^{3+}$  affords the cobalt(III) sarcophagine(sar)-type complexes with stabilized diamine chelate ring conformations; we describe the synthesis, resolution and tentative identity of the one *fac* and three *mer* geometrical isomers of the *lel<sub>2</sub>ob* cages, as well as the subsequent conversion to their *ob<sub>2</sub>lel* diastereoisomers.

The ambiguity introduced by the conformational flexibility of the  $[M(X)_{2}$ -sar]<sup>n+</sup> complexes (Figure 1; X = apical cap substituents, sar = 3,6,10,13,16,19-hexa-azabicylo[6.6.6]eicosane)<sup>1</sup> necessitates formation of cage complexes with stabilized diamine chelate ring conformations.<sup>2.3</sup>  $[Co(R/S-pn)_3]^{3+}$ (pn = propane-1,2-diamine) provides a unique series of twelve enantiomeric pairs ( $\Delta$  and  $\Lambda$ ) of diastereoisomers with parallel and oblique conformational variations ( $lel_3$ ,  $lel_2ob$ ,  $ob_2lel$ , and  $ob_3$ ),<sup>4</sup> by virtue of the pn methyl substituent favouring an equatorial orientation on the five-membered chelate rings.<sup>5</sup> A statistical distribution results in three *mer*idional and one *fac*ial geometrical isomers for each of the conformational variants.<sup>6,7</sup> These *mer* isomers are degenerate for (pseudo-)  $C_3$  symmetry, which provides a statistical predominance of 3 *mer*: 1 *fac* in the *lel*<sub>3</sub> and *ob*<sub>3</sub> systems.

Recently, one of us synthesised and resolved the  $lel_3$ - and  $ob_3$ -[Co(NH<sub>2</sub>)<sub>2</sub>-pnsar]<sup>3+</sup> cage diastereoisomers (Figure 2;



Figure 1.  $[M(X)_2$ -sar $]^{n+}$  complex.

fac-pnsar = 1,8-diamino-4,12,17-trimethylsar; *mer*-pnsar = 1,8-diamino-4,11,17-trimethylsar).<sup>8,9</sup> The  $lel_3$  cage isomers have surprisingly different characteristics compared to those of their  $ob_3$  diastereoisomers, but their respective geometrical fac and *mer* isomers have virtually identical properties.<sup>8</sup> We now report the syntheses of the  $lel_3$ ,  $lel_2ob$ , and  $ob_2lel$  cage complexes in a one-pot reaction from an equilibrium mixture of  $[Co(R/S-pn)_3]^{3+}$  substrate  $(lel_3 = -36\%, lel_2ob = 42\%, ob_2lel = 18\%$  and  $ob_3 = 4\%$ )<sup>7</sup> using formaldehyde/nitromethane under basic conditions,<sup>1</sup> to give a mixture of the  $[Co(NO_2)_2-pnsar]^{3+}$  diastereoisomers. After reducing the cap nitro



Figure 2. Interconversion of  $lel_3$ - and  $ob_3$ - $[Co(NH_2)_2$ -pnsar $]^{3+}$  complexes. Fac when A = methyl above and B = H below; Mer when B = methyl below, A = H above (H atoms are excluded for clarity and  $\bullet$  = N).



Scheme 1. Top view of the conversion of  $\Lambda(S_2R)$ -lel<sub>2</sub>ob isomers to their  $\Delta(S_2R)$ -ob<sub>2</sub>lel diastereoisomers. A = methyl above, B = methyl below.

substituents to amines with Zn/HCl,1 the isomer mixture was initially resolved on an SP Sephadex C-25 cation exchange column using 0.20 м K<sub>2</sub>SO<sub>4</sub> eluent; two major front orange bands and one minor back orange band were isolated. The two major bands were further resolved on a one-metre SP Sephadex C-25 column using 0.25 м Na<sub>3</sub>PO<sub>4</sub> eluent. The first band ( $\sim 23\%$  yield) was resolved into two products and identified by <sup>13</sup>C n.m.r. as the mer and fac isomers of the  $lel_3$ -[Co(NH<sub>2</sub>)<sub>2</sub>-pnsar]<sup>3+</sup> cation.<sup>8</sup> The second band (~11%) yield) was similarly resolved into four products by continuous recycling on the column. These have been identified as the one fac and three mer lel<sub>2</sub>ob isomers, indicative from their composite <sup>13</sup>C spectral features of the pure  $lel_3$  and  $ob_3$  cages. The four very similar  $lel_2ob$  <sup>13</sup>C spectra each show the expected 17 carbon atoms.<sup>†</sup> The lower yield of the *lel<sub>2</sub>ob* cage is due to the increased ob character of the starting substrate. This rationale follows from the fact that  $lel_3$ -[Co(pn)<sub>3</sub>]<sup>3+</sup> encapsulates with  $\sim 70\%$  success while the  $ob_3$  substrate fails to form a cage.8

The tentative identification of these  $lel_2ob$  isomers has been rationalized by their elution order relative to the elution behaviour of the  $lel_3$  and  $ob_3$  cages. The  $lel_3$  cages elute well ahead of their  $ob_3$  diastereoisomers,<sup>8</sup> indicative of the more open pseudo- $C_2$  lel faces which have the protons on the co-ordinated amines better disposed for hydrogen-bonding with counter-ions. Furthermore, in both  $lel_3$  and  $ob_3$  cases the *mer* isomers elute marginally faster than their *fac* partners, indicative of the more accessible faces that eventuate in the respective *mer* situations. Hence the slowest moving  $lel_2ob$  isomer (4) must be fac (Scheme 1). The fastest moving isomer (1) must have the open  $mer-lel_2$  face, while isomer (3), with the closed  $mer-lel_2$  face, eluted more slowly. Isomer (2), with the fac-lel<sub>2</sub> face, co-eluted with (1) for five cycles on the column before a separation became obvious. This identification is further corroborated by the unequal amounts of the *lel<sub>2</sub>ob* cage isomers that were formed, despite the probability that the  $lel_2ob$ -[Co(pn)<sub>3</sub>]<sup>3+</sup> substrate isomers form in approximately equal amounts under equilibrium conditions.<sup>7</sup> The yields indicate a pattern of reactivity which is a function of the *lel* and *ob* character of the initial substrate as well as the steric influence of the pn methyl substituents. The fac-lel<sub>2</sub>ob isomer (4), with the one trigonal  $N_3$  face crowded by one *ob* and two *lel* methyl substituents, gave the lowest relative yield ( $\sim 5\%$ ). Transposing the ob methyl to the opposite trigonal N<sub>3</sub> face, producing the mer isomer (2) with both the top and bottom faces least hindered, gave the highest relative yield ( $\sim 43\%$ ). The other two intermediate cases, where the *ob* methyl shares an N<sub>3</sub> trigonal face with one of the *lel* methyl substituents, gave intermediate relative yields of isomers (1) ( $\sim$ 24%) and  $(3) (\sim 28\%).$ 

The <sup>13</sup>C n.m.r. spectrum of the third minor band (<1% yield) from the initial reaction indicated that it is mainly a single isomer of  $ob_2lel$  cage which has the required composite spectral features of the  $ob_3$  and  $lel_3$  diastereoisomers.<sup>†</sup> This drastic decrease in capping ability is due to the increased ob character of the substrate and an alternative stategy was adopted to synthesise and identify the  $ob_2lel$  cage isomers. This involved the same procedure developed for the synthesis of the  $ob_3$  cages.<sup>8</sup> As the  $ob_3$  substrates did not encapsulate, the  $ob_3$  cages were eventually synthesised by removing Co<sup>III</sup> from the  $lel_3$  cage complexes in concentrated NaCN and then reinserting a labile form of Co<sup>III</sup> into the free cage ligand, *via trans*-[Co(py)<sub>4</sub>Cl<sub>2</sub>]Cl (py = pyridine) in 2-methoxyethanol, effectively isomerizing, for example, the  $\Lambda(S_3)$ -lel\_3 cage to its  $\Lambda(S_3)$ -ob\_3 diastereoisomer (Figure 2). We performed this

<sup>&</sup>lt;sup>+</sup> <sup>13</sup>C N.m.r. signals, assigned by APT and DEPT, occur in the following ranges (D<sub>2</sub>O with external reference dioxane =  $\delta$  67.3): pn-CH<sub>3</sub>, *lel* =  $\delta$  13.6–14.4, *ob* =  $\delta$  12.4–12.7; pn-CH<sub>2</sub>, *lel* =  $\delta$  60.7–61.9, *ob* =  $\delta$  55.2–57.2; pn-CH, *lel* =  $\delta$  63.8–65.2, *ob* =  $\delta$  57.4–59.3; Cap-CH<sub>2</sub>,  $\delta$  45.2 (adjacent to *ob* pn-CH<sub>3</sub>) to 54.7; Cap quaternary *C*, 55.6–58.2.

interconversion on the three *mer* isomers of the *lel*<sub>2</sub>*ob* cages, in effect transforming, for example, the  $\Lambda(S_2, R)$ -*lel*<sub>2</sub>*ob* cages to their  $\Delta(S_2, R)$ -*ob*<sub>2</sub>*lel* diastereoisomers in overall conversions of ~10% (Scheme 1). The conversion product from *lel*<sub>2</sub>*ob* (1) has the same <sup>13</sup>C n.m.r. spectrum as that of the *ob*<sub>2</sub>*lel* cage (1\*) isolated from the original one-pot template synthesis, which is consistent with the reactivity discussed above.

A surprising feature of this cage system occurs in the visible spectral regions. The first visible transition, of  ${}^{1}A_{1g} \rightarrow {}^{1}T_{1g}$ origin, occurs at 476 nm ( $\epsilon \sim 150 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ ) for the  $lel_3$ -[Co(NH<sub>2</sub>)<sub>2</sub>-pnsar]<sup>3+</sup> cages and at 450 nm ( $\varepsilon$  ~85  $dm^3 mol^{-1} cm^{-1}$ ) for the  $ob_3$  cages. Although one might expect the *lel<sub>2</sub>ob* and the *ob<sub>2</sub>lel* cages to have transitions at intermediate wavelengths, the lel<sub>2</sub>ob transition occurs at 484 nm ( $\varepsilon \sim 160 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ ) and the *ob\_2lel* at 475 nm ( $\varepsilon \sim 200$ dm<sup>3</sup> mol<sup>-1</sup> cm<sup>-1</sup>), both close to the *lel*<sub>3</sub> values. The respective fac/mer isomerism has minimal effect on these visible transitions. Molecular mechanics calculations show that the  $ob_3$ cages have close to octahedral CoN<sub>6</sub> core symmetry and the shortest Co-N bond lengths,8 and this reflects in the lower extinction coefficient and shorter wavelength. The lel3 cages undergo some trigonal twist (the twist angle,  $^{10} \phi \sim 56^{\circ}$ ) imparting  $C_3$  symmetry to the CoN<sub>6</sub> chromophore. The rise in extinction coefficients for the lel<sub>2</sub>ob and ob<sub>2</sub>lel cage isomers reflect a further loss of CoN<sub>6</sub> inner co-ordination-sphere symmetry.

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## References

- 1 R. J. Geue, T. W. Hambley, J. M. Harrowfield, A. M. Sargeson, and M. R. Snow, J. Am. Chem. Soc., 1984, 106, 5478.
- 2 I. I. Creaser, R. J. Geue, J. M. Harrowfield, A. J. Herlt, A. M. Sargeson, M. R. Snow, and J. Sprinborg, J. Am. Chem. Soc., 1982, 104, 6016.
- 3 R. J. Geue, M. G. Mc Carthy, and A. M. Sargeson, J. Am. Chem. Soc., 1984, 106, 8282.
- 4 IUPAC Commission on the Nomenclature of Inorganic Chemistry, Inorg. Chem., 1970, 9, 1; 1977, 16, 233.
- 5 E. J. Corey and J. C. Bailar, J. Am. Chem. Soc., 1959, 81, 290.
- 6 F. P. Dwyer, F. L. Garvan, and A. Schulman, J. Am. Chem. Soc., 1959, 81, 290.
- 7 S. E. Harnung, S. Kallesoe, A. M. Sargeson, and C. E. Schaffer, Acta Chem. Scand., 1974, 28, 385.
- 8 A. J. Hendry, Ph.D. Thesis, Australian National University, 1986.
- 9 A. J. Hendry, R. J. Geue, and A. M. Sargeson, J. Chem. Soc., Chem. Commun., submitted for publication.
- 10 P. Comba, A. M. Sargeson, L. M. Engelhardt, J. M. Harrowfield, A. H. White, E. Horn, and M. R. Snow, *Inorg. Chem.*, 1985, 24, 2325.